

As₂Na₆O₈ (s) 3Na ₂ O · As ₂ O ₅ (s)	Sodium Arsenate gamma	As₂Na₆O₈ (s) 3Na ₂ O · As ₂ O ₅ (s)
$\Delta H_{733}^0 = -2886.3 \text{ kJ}\cdot\text{mol}^{-1}$ [4] $C_p^0 = 271.12 + 234.3 \cdot 10^{-3} \cdot T \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ (733 ... 1000 K) [4]		$S_{733}^0 = 750.7 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ [4]
As₂Ni₃O₈ (s) 3NiO · As ₂ O ₅ (s)	Nickel Arsenate	As₂Ni₃O₈ (s) 3NiO · As ₂ O ₅ (s)
$\Delta H_{298}^0 = -1849.2 \text{ kJ}\cdot\text{mol}^{-1}$ [2] $C_p^0 = 265.4 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ (298 K) [2]		$S_{298}^0 = 344.8 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ [2]
As₂Ni₅ (s) Ni ₅ As ₂ (s)	Nickel Arsenide	As₂Ni₅ (s) Ni ₅ As ₂ (s)
$\Delta H_{298}^0 = -251.1 \text{ kJ}\cdot\text{mol}^{-1}$ [2] $C_p^0 = 216 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ (298 K) [2]		$S_{298}^0 = 190.6 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ [2]
As₂O₃ (s)	Arsenic(III) Oxide	As₂O₃ (s)
mp = 588 K (315 °C) $\Delta H_{298}^0 = -655 \text{ kJ}\cdot\text{mol}^{-1}$ [4] $C_p^0 = 93.71 + 58.48 \cdot 10^{-3} \cdot T - 1.26 \cdot 10^6 \cdot T^{-2} \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ (298 ... 607 K) [4] $\lg(p, K) = -6.49 \cdot 10^3 \cdot T^{-1} - 3.69 \cdot \lg(T) + 20.46$ (400 ... 607 K) [4] {Reaction: evaporation as As ₄ O ₆ (g)}		$S_{298}^0 = 113.3 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ [4]
As₂O₅ (s)	Arsenic(V) Oxide	As₂O₅ (s)
$\Delta H_{298}^0 = -924.9 \text{ kJ}\cdot\text{mol}^{-1}$ [2] $C_p^0 = 112.21 + 82.94 \cdot 10^{-3} \cdot T - 1.81 \cdot 10^6 \cdot T^{-2} \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ (298 ... 1084 K) [4]		$S_{298}^0 = 105.4 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ [2]
As₂O₈Pb₃ (s) 3PbO · As ₂ O ₅ (s)	Lead Arsenate	As₂O₈Pb₃ (s) 3PbO · As ₂ O ₅ (s)
$\Delta H_{298}^0 = -1780.2 \text{ kJ}\cdot\text{mol}^{-1}$ [2] $C_p^0 = 258 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ (298 K) [2]		$S_{298}^0 = 324.6 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ [2]