

2.17



Вещь	$\frac{\text{Дж}}{\text{моль}}$ ΔH_f°	$\Delta H_{\text{ком}}$
Al_2O_3	50,94	-1676
SO_2	248,1	-296,9
O_2	205,09	0
$\text{Al}_2(\text{SO}_4)_3$	(?)	(?)

$$\Delta H_f^\circ = \Delta H(\text{Al}_2(\text{SO}_4)_3) - (\Delta H_f^\circ(\text{Al}_2\text{O}_3) \cdot 2 + 6\Delta H_f^\circ(\text{SO}_2) + 3 \cdot \Delta H_f^\circ(\text{O}_2))$$

0

в метрической системе единиц

$$\Delta H_f^\circ = \Delta H(\text{Al}_2(\text{SO}_4)_3) - (2 \cdot (-1676) + 6 \cdot (-296,9))$$

$$\Delta H_f^\circ = \Delta H(\text{Al}_2(\text{SO}_4)_3) - (-3352 - 1781,4)$$

$$\Delta H_f^\circ = \Delta H(\text{Al}_2(\text{SO}_4)_3) + 5133,4 \text{ Дж/моль}$$

$$\Delta S_{p-n} = \Delta S(\text{Al}_2(\text{SO}_4)_3) - (2\Delta S(\text{Al}_2\text{O}_3) + 6\Delta S(\text{SO}_2) + 3\Delta S(\text{O}_2))$$

$$= \Delta S(\text{Al}_2(\text{SO}_4)_3) - (2 \cdot 50,94 + 6 \cdot (248,1) + 3 \cdot 205,09)$$

$$= \Delta S(\text{Al}_2(\text{SO}_4)_3) - (101,88 + 1488,6 + 615,12)$$

$$= \Delta S(\text{Al}_2(\text{SO}_4)_3) - 2205,6 \text{ Дж/моль}$$

